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oxidation number of

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the underlined element: KMnO_4 . Since K is an alkali metal, its charge must be $1+$. Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$. If $1+$ and $8-$ are added together, we get $7-$. In order for the compound to be neutral, the Mn must be $7+$.

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$1C_6H_{14} + 2O_2 \rightarrow$
 $1_2CO_2 + MH_2O$. 5.

$Ca_3(PO_4)_2 + ZH_2SO_4$

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water. How to identify?

... The solubility rules came out of page 49 of The Ultimate Chemical Equations Handbook (student edition) ...

you'll have your answer. As water is a polar molecule, only other polar molecules will be soluble. 0 0 3.

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i.e.,

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German names. Ag Sn
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(argentum) (stannum)
(plumbum) (wolfram)
iron sodium silver tin
lead tungsten (cuprum)
(kalium) (hydrargyrum)
(stibium) (aurium)
copper potassium
mercury antimony gold
Symbols are used as a
sort of shorthand in
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elements.

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(ii) from: Rodney
Schmitz (2007) Moore
County Schools, NC For
each of the following
eight reactions, in part
(i) write a BALANCED
equation and in part (ii)
answer the question
about the reaction.

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Answers for the
questions in Chapter 8:
ROUND 3 Exercise 8-1:
Using the activity
series, predict and
balance the following
single replacement
reactions. Use

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abbreviations to indicate the appropriate phase of reactants and products where possible. Note: Not all of the reactions will occur. For those that do not, write no reaction. I.

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